Chemistry Equation Writing Activity

For each equation: 1) write the full equation (including states)

2) balance the equation

3) name the equation type

4) write the complete ionic equation

5) write the net ionic equation \*\*If no reaction occurs, write NR on product side

Zinc nitrate (aq) and sodium phosphate (aq) yields

Mercury (I) nitrate (aq) and copper (II) chloride (aq) yields

HClO4(aq) + NaOH (aq) →

NiCl2(aq) + Na3PO4(aq) →

CuSO4(aq) + NaOH(aq)→

NaCl(aq) + Ca(NO3)2(aq) →

Ba(NO3)2(aq) + Na2CO3(aq)→

HCl(aq) + NaOH(aq) →

KOH(aq) + HNO3(aq)→

Ba(OH)2(aq) + NaOH(aq)→

Fe(NO3)3(aq) + KOH(aq)→

AgNO3(aq) + NH4Cl(aq)→

Na2SO4(aq) + BaCl2(aq) →

HBr(aq) + Ba(OH)2(aq) →

(NH4)2S(aq) + Co(NO3)2(aq) →

Nitric acid and strontium hydroxide yield

Lead (II) acetate and sodium chloride yield